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ΔH (IF) $-1 = -840$ kJ mol A. -190 B. -95 C. $+6$ D. $+95$ -1 7 2 5 7 -1 5 Which is a correct definition of lattice enthalpy? A. It is the enthalpy change that occurs when an electron is removed from 1 mol of gaseous atoms. B. It is the enthalpy change that occurs when 1 mol of a compound is formed from its elements.

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1 (SL): 45 minutes. 30 multiple-choice questions on core, about 15 of which are common with HL. The questions on paper 1 test assessment objectives 1, 2 and 3. The use of calculators is not permitted.

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