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Chapter 11 Small-Scale Lab Section 11.3 Precipitation Reactions: Formation of Solids, page 345 Analysis 1. a. $\text{Na}_2\text{CO}_3 + 2\text{AgNO}_3 \rightarrow 2\text{NaNO}_3 + \text{Ag}_2\text{CO}_3(\text{s})$ b. $2\text{Na}_3\text{PO}_4 + 3\text{Pb}(\text{NO}_3)_2 \rightarrow 6\text{NaNO}_3 + \text{Pb}_3(\text{PO}_4)_2(\text{s})$ 2. Sodium hydroxide reacts with calcium chloride to form sodium chloride and solid calcium hydroxide. 3. Mixings d, n, and o all gave no visible

Chapter 11 Small-Scale Lab

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